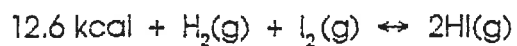
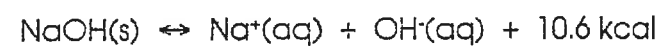


LE CHATELIER'S PRINCIPLE CONTINUED

Name _____



Stress	Equilibrium Shift	[H ₂]	[I ₂]	[HI]
1. Add H ₂	right	_____	decreases	increases
2. Add I ₂			_____	
3. Add HI				_____
4. Remove H ₂		_____		
5. Remove I ₂			_____	
6. Remove HI				_____
7. Increase Temperature				
8. Decrease Temperature				
9. Increase Pressure				
10. Decrease Pressure				



Stress	Equilibrium Shift	Amount NaOH(s)	[Na ⁺]	[OH ⁻]
1. Add NaOH(s)		_____		
2. Add NaCl (Adds Na ⁺)			_____	
3. Add KOH (Adds OH ⁻)				_____
4. Add H ⁺ (Removes OH ⁻)				_____
5. Increase Temperature				
6. Decrease Temperature				
7. Increase Pressure				
8. Decrease Pressure				

LE CHATELIER'S PRINCIPLE

Name _____

5

Le Chatelier's Principle states that when a system at equilibrium is subjected to a stress, the system will shift its equilibrium point in order to relieve the stress.

Complete the following chart by writing left, right or none for equilibrium shift, and decreases, increases or remains the same for the concentrations of reactants and products, and for the value of K.



Stress	Equilibrium Shift	[N ₂]	[H ₂]	[NH ₃]
1. Add N ₂	right	_____	decreases	increases
2. Add H ₂			_____	
3. Add NH ₃				_____
4. Remove N ₂		_____		
5. Remove H ₂			_____	
6. Remove NH ₃				_____
7. Increase Temperature				
8. Decrease Temperature				
9. Increase Pressure				
10. Decrease Pressure				