

# HYDROLYSIS OF SALTS

Name \_\_\_\_\_

Salt solutions may be acidic, basic or neutral, depending on the original acid and base that formed the salt.

Strong Acid + Strong Base  $\rightarrow$  Neutral Salt

Strong Acid + Weak Base  $\rightarrow$  Acidic Salt

Weak Acid + Strong Base  $\rightarrow$  Basic Salt

A weak acid and a weak base will produce any type of solution depending on the relative strengths of the acid and base involved.

Complete the table below for each of the following salts.

Salt	Parent Acid	Parent Base	Type of Solution
1. KCl			
2. $\text{NH}_4\text{NO}_3$			
3. $\text{Na}_3\text{PO}_4$			
4. $\text{CaSO}_4$			
5. $\text{AlBr}_3$			
6. $\text{CuI}_2$			
7. $\text{MgF}_2$			
8. $\text{NaNO}_3$			
9. $\text{LiC}_2\text{H}_3\text{O}_2$			
10. $\text{ZnCl}_2$			
11. $\text{SrSO}_4$			
12. $\text{Ba}_3(\text{PO}_4)_2$			